

# Cloudy's Rainy Day

By Derrick G. Hamlin



Illustrated by Rhianon Dameron

# Rainy Days English Edition

**K Morrison**



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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H<sub>2</sub>S and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>S from the equation  
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} \times 18 \text{ g H}_2\text{O} = 149.4 \text{ g H}_2\text{O}$  is your theoretical yield so percent yield  
 $\frac{137.1 \text{ g H}_2\text{O}}{149.4 \text{ g H}_2\text{O}} \times 100 = 91.77\%$  problems 6 & 8 are similar to problem 5

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web 1.csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH<sub>3</sub> can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  
 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of O<sub>2</sub> produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  
 $0.2162 \text{ mol O}_2 \times \frac{32 \text{ g}}{1 \text{ mol}} = 6.92 \text{ g O}_2$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency =  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction  
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub>  
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  c determine the percent yield  
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$  theoretical 302.3 g

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H<sub>2</sub> are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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