

# 14.1

## HYDROGEN IONS AND ACIDITY

### Section Overview

#### Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydronium and hydroxide ion concentrations, pH, and pOH, and vice versa, for strong acids and bases.
3. Calculate the pH of weak acids.

#### Vocabulary

- |                   |                             |                     |
|-------------------|-----------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. strong acid/base |
| 4. hydroxide ion  | 5. hydronium ion            | 6. weak acid/base   |
|                   | 7. acid solution            | 8. pH               |

#### Key Equations

- |  |  |
|--|--|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |
| 3. $\text{pOH} = -\log[\text{OH}^-]$                                 | 4. $\text{pH} + \text{pOH} = 14$             |

#### Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the \_\_\_\_\_ of a weak hydroxide ion and a hydroxide ion? \_\_\_\_\_ of the conjugate acid-base pair.
2. \_\_\_\_\_ of a strong acid-base pair \_\_\_\_\_ water.
3. \_\_\_\_\_ of a weak acid-base pair \_\_\_\_\_ water.
4. \_\_\_\_\_ of a conjugate acid-base pair \_\_\_\_\_ the other ion.
5. \_\_\_\_\_ of a conjugate acid-base pair \_\_\_\_\_ the original acid-base pair.
6. \_\_\_\_\_ of a weak acid-base pair \_\_\_\_\_ water and \_\_\_\_\_.
7. \_\_\_\_\_ of a conjugate acid-base pair \_\_\_\_\_ water.
8. \_\_\_\_\_ of a conjugate acid-base pair \_\_\_\_\_ water.

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