

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide ion concentrations, molar ratios, and pH values for weak acid-conjugate base systems.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                             |                  |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion  | 5. molar ratio              | 6. hydroxide ion |
|                   | 7. weak acid                | 8. pH            |

## Key Equations

- |  |   |
|--|---|
| 1. $\text{pH} = -\log[\text{H}^+]$ (Eq. 14.1)        | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.4)  |

## Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that are based on the following data:  $K_a(\text{HClO}_2) = 1.1 \times 10^{-2}$ .

1. What products are **1** formed when hydrochlorous acid ( $\text{HClO}_2$ ) ionizes in water? The concentrations of these species produced in 0.100 M  $\text{HClO}_2$  solution are **2**  $\text{M}$ .
2. Hydrochlorous acid ionizes in water **3**  $\text{M}$  in a 0.100 M solution. **4**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution. **5**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution. **6**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution.
3. **7**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution. **8**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution. **9**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution. **10**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution.
4. **11**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution. **12**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution. **13**  $\text{M}$  hydrochlorous acid ionizes in the 0.100 M solution.

1.  $\text{HClO}_2$
1.  $1.1 \times 10^{-2}$
- 0.100
- 0.0011
- 0.0011
- 0.0011
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# 19 2 Hydrogen Ions From Water Answers

**Wenbin Ji**



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