

14.5

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH and pOH and the associated logarithmic functions.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. hydronium ion | 6. logarithmic function |
| | 7. weak solution | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{pH} = -\log_{10} [\text{H}^+]$ (Eq. 14.1) | 2. $\text{pOH} = -\log_{10} [\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.4) |

Part A: Completion

Use the logarithmic function to calculate the missing quantity of the conjugate acid-base pair represented in the partial acid-base ionization reactions below. Complete as needed.

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JR Anderson



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