



# 79 Wildlife Patterns for the Scroll Saw



Frank Puzgall



A Schiffer Book for Woodworkers

# 79 Wildlife Patterns For The Scroll Saw

**JS Bruner**



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2 p 1 b 1 s 2 2 s 1 2 p 1 1 s 2 2 p 2 c 1 s 2 2 p 2 1 s 2 1 p 2 d 1 s 2 1 p 2

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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s 2 2 s 1

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web the electron removed during the ionization of beryllium he 2s 2 is an s electron whereas the electron removed during the ionization of boron he 2s 2 2p 1 is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s 2 ne 10 neon 1s 2 2s 2 2p 6

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potassium element has 19 electrons its electron configuration is  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$  or simply put as  $4s^1$  what is the total number of electrons that can be filled in s p and d subshells

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web by building up from hydrogen this table can be used to determine the electron configuration for any atom on the periodic table we will now construct the ground state electron configuration and orbital diagram for a selection of atoms in the first and second periods of the periodic table

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