

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from K_w and K_a or K_b and the concentration of a weak acid or base.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. K_a | 6. K_b |
| | 7. weak base | |

Key Equations

- | | |
|--|---|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ | 4. $\text{pOH} = -\log[\text{OH}^-]$ |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions. Show calculations for all problems. Answers should include units unless otherwise specified.

1. What products are formed when hydrochloric acid and methanamine react? Write the chemical formula of these acids, products, or hydroxides and explain.
2. Hydrochloric acid ionizes 100% in water. Calculate the pH of a 0.010 M solution.
3. Calculate the pH of a 0.010 M solution of a weak acid with $K_a = 1.0 \times 10^{-4}$.
4. Calculate the pH of a 0.010 M solution of a weak base with $K_b = 1.0 \times 10^{-4}$.
5. Calculate the pH of a 0.010 M solution of a weak acid with $K_a = 1.0 \times 10^{-4}$.
6. Calculate the pH of a 0.010 M solution of a weak base with $K_b = 1.0 \times 10^{-4}$.
7. Calculate the pH of a 0.010 M solution of a weak acid with $K_a = 1.0 \times 10^{-4}$.
8. Calculate the pH of a 0.010 M solution of a weak base with $K_b = 1.0 \times 10^{-4}$.

9. Calculate the pH of a 0.010 M solution of a weak acid with $K_a = 1.0 \times 10^{-4}$.
10. Calculate the pH of a 0.010 M solution of a weak base with $K_b = 1.0 \times 10^{-4}$.
11. Calculate the pH of a 0.010 M solution of a weak acid with $K_a = 1.0 \times 10^{-4}$.
12. Calculate the pH of a 0.010 M solution of a weak base with $K_b = 1.0 \times 10^{-4}$.
13. Calculate the pH of a 0.010 M solution of a weak acid with $K_a = 1.0 \times 10^{-4}$.
14. Calculate the pH of a 0.010 M solution of a weak base with $K_b = 1.0 \times 10^{-4}$.
15. Calculate the pH of a 0.010 M solution of a weak acid with $K_a = 1.0 \times 10^{-4}$.
16. Calculate the pH of a 0.010 M solution of a weak base with $K_b = 1.0 \times 10^{-4}$.

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