

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide ion concentrations, molar ratios, and pH values for weak acid-conjugate base systems.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                             |                  |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion  | 5. weak acid                | 6. hydroxide ion |
|                   | 7. weak base                | 8. pH            |

## Key Equations

- |  |   |
|--|---|
| 1. $\text{pH} = -\log[\text{H}^+]$ (Eq. 14.1)        | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.4)  |

## Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the system. Also indicate which species acts as the conjugate acid.

1.  $\text{H}_2\text{O}$  and  $\text{OH}^-$  (conjugate acid-base pair)
2.  $\text{H}_2\text{O}$  and  $\text{H}_3\text{O}^+$  (conjugate acid-base pair)
3.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
4.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
5.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
6.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
7.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
8.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
9.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
10.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
11.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
12.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
13.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
14.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
15.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
16.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
17.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
18.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
19.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
20.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)

## Part B: Multiple Choice

1.  $\text{H}_2\text{O}$  and  $\text{OH}^-$  (conjugate acid-base pair)
2.  $\text{H}_2\text{O}$  and  $\text{H}_3\text{O}^+$  (conjugate acid-base pair)
3.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
4.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
5.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
6.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
7.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
8.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
9.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
10.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
11.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
12.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
13.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
14.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
15.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
16.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
17.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
18.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
19.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)
20.  $\text{H}_2\text{O}$  and  $\text{H}^+$  (conjugate acid-base pair)

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**M Walker**



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