

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH⁻] and [H⁺].
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|---|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H ₃ O ⁺) | 6. autoionization |
| | 7. acid solution | 8. pH |

Key Equations

- | | |
|--|------------------------|
| 1. $[H^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H^+]$ |
| 3. $pH = -\log[H^+]$ | 4. $pOH = -\log[OH^-]$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the _____ of a weak hydroxide ion and a hydronium ion? _____ of the autoionization of water which produces an OH⁻ ion and a H₃O⁺ ion.
2. Hydrochloric acid ionizes from _____ to a Cl⁻ ion and a H⁺ ion.
3. _____ ionization is observed with the autoionization of water.
4. _____ ionization _____ of water _____ the autoionization of water.
5. _____ ionization is used to describe the autoionization of water.
6. _____ ionization _____ of water _____ the autoionization of water.
7. _____ ionization is used to describe the autoionization of water.

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