

EARTH SCIENCE

INTRODUCTION

- COMPONENTS OF STUDY**
1. **Geology:** Both physical & historical
 2. **Meteorology:** Weather & climate
 3. **Oceanography:** The oceans
 4. **Astronomy:** What is external to the atmosphere
- THE FOUR SPHERES (See The Water Cycle)**
1. **Lithosphere:** The solid earth
 2. **Hydrosphere:** The liquid earth
 3. **Atmosphere:** The gaseous earth
 4. **Biosphere:** Where life exists
- CONCERNS IN OUR PHYSICAL ENVIRONMENT**
1. **Resources:** Water, soil, metallic & non-metallic elements, sunlight, wind, etc.
 2. **Per-capita use of metallic and non-metallic elements is 11 tons per year**
 3. **Renewable vs. non-renewable resources**
 4. **Environmental and population problems**

MINERALS

- A **mineral** is a naturally occurring, inorganic, solid material with a definite atomic structure.
 - An **ore** is a useful metallic mineral.
- A. Atom and Forces**
1. **Atom,** in ancient Greek, means "no-cut," i.e., the Ancients believed the atom to be the smallest particle of matter.
 2. Atoms are composed of neutrons, protons and electrons.
 - a. **Atomic number:** Number of protons in nucleus
 - b. **Atomic weight:** Average weight of an atom
 - c. **Isotope:** Forms of an element with identical atomic numbers, but different numbers of neutrons in the nucleus
 3. **Force:** Strong, weak, electromagnetic and gravity
- B. Properties of Minerals**
1. **Color:** Appearance or quality of light from surface
 2. **Streak:** Color of mineral in powdered form
 3. **Hardness:** Moh's Scale (See figure below)
 4. **Scratch Test:** Higher numbered materials can scratch lower numbered materials
 5. **Shape** is determined by cleavage, crystal form and fracture
 6. **Specific Gravity**
 - a. The ratio of the weight of a mineral to the weight of an equal volume of water
 - b. **Density of water = 1 gm/cm³ = 1 gm/ml.**
 i.e., lead = 7.7, aluminum = 2.7, osmium = 22

MOH'S SCALE		
Hardness	10	MINERAL
	9	Diamond
	8	Corundum
	8	Topaz
	7	Quartz
	6	Feldspar
	5	Apatite
	4	Fluorite
	3	Calcite
	2	Gypsum
	1	Talc

- C. Mineral Groups**
1. **Silicates:** Minerals with silicon and oxygen
 - a. **Silicon** is a semi-metal which forms a pyramid-shaped structure with oxygen.
 - b. **Oxygen** forms inside single tetrahedrons (silvins), chains (single) sheets (micas), 3-D structures (feldspars and quartz)
 2. **Non-Silicates**
 - a. Make up one-fourth of continental crust
 - b. **Carbonates:** Minerals with carbon and oxygen including calcite, from which we procure limestone (roads) and marble (decorative slabs)
 - c. **Oxides:** Oxygen based solids, i.e., iron sulfides (S), sulfates (SO₄), halides (Cl, F)
 - d. **Halide:** Mineral form of salt
 - e. **Gypsum:** Plaster, calcium
 - f. **Native metals:** Iron, zinc, gold, silver, nickel

ROCK TYPES



- IGNEOUS ROCKS**
- A. Formation**
1. **Extrusive igneous rocks** are formed in volcanic eruptions at surface.
 2. **Intrusive (or plutonic) igneous rocks** are produced within the earth.
 3. **Magma**
 - a. Consists mainly of silicate materials
 - b. Has gases, such as water vapor
 - c. Originates from as deep as 200 km
 - d. Differs in rate of cooling, chemical composition, and amount of gases
 - e. Produces intrusive (plutonic) igneous rocks
 4. **Lava**
 - a. Similar to magma
 - b. Most gaseous elements have escaped
 - c. Produces extrusive (volcanic) igneous rocks
- B. Structure**
1. **Crystallization (ordered pattern of ions)**
 - a. The rate of cooling strongly affects crystal size
 - b. Slow cooling = large crystals
 - c. Quick cooling = small crystals
 2. **Amorphous (unordered pattern of ions)**
 - a. Occurs with instant cooling (quenching)
 - b. Produces rocks referred to as glass
- C. Classifying Igneous Rocks**
1. **By texture, size and arrangement of interlocking crystals**
 - a. **fine-grained**
 - i. Form at surface or within upper crust
 - ii. Openings left by gas bubbles are called vesicles
 - b. **coarse-grained:** Formed deep within the interior
 - c. **porphyritic:** Large crystals imbedded in matrix of smaller crystals
 2. **By mineral composition**
 - a. Depends on composition of magma and cooling
 - b. Same magma can create rocks of varying mineral content.
 - c. As magma cools, certain minerals crystallize first (at very high temperatures).

- METAMORPHIC ROCKS**
- A. Formation**
1. **Metamorphic:** To change form
 2. **Regional metamorphism:** Material under intense stresses, high temperatures
 3. **Contact (thermal) metamorphism:** Changes caused by proximity to magma
 4. **Low grade metamorphism:** Slate becomes shale
 5. **High grade metamorphism:** Fossils become rock
 6. **Cliffs form** during mountain building
 7. Most metamorphic rocks are harder than sedimentary rocks
 8. **Metamorphism changes texture**
 9. **Low-grade metamorphism** makes material more compact, more dense
 10. **Foliated texture:** Particles of material are brought into line with one another
- B. Agents of Metamorphism**
1. **Heat**
 - a. Most important agent
 - b. Provides energy for chemical reactions
 - c. For example: Clay recrystallizes into a mineral at great temperature
 2. **Pressure**
 3. **Chemical Activity:** Most common chemical agent in water
- C. Types of Metamorphic Rocks**
1. **Gneiss:** Fine grained, dark flinty rock with randomly arranged minerals
 2. **Slate:** Fine grained, often gray, foliated rock split easily along cleavage, planes of mica flakes aligned by pressure
 3. **Marble:** Granular or sugary-textured rock formed from limestone
 4. **Phyllite:** Silky, foliated rock more coarsely grained than slate
 5. **Schist:** foliated rock, more coarsely grained and of higher metamorphic grade than phyllite
 6. **Gneiss:** Foliated, banded rock of the highest metamorphic grade
 7. **Quartzite:** Very hard, granular quartz rock, formed from sandstone

- BOWEN'S REACTION SERIES**
- If a mineral, which has already formed, remains in the magma, it will react with the remaining magma to produce the next mineral in the sequence. For example, *olivine forms first; olivine then reacts with remaining magma to form pyroxene.*
- A. Right Side of the Bowen Series**
1. Calcium rich parts of the magma form small crystals of feldspar.
 2. These react with sodium in the magma to become more and more sodium rich.
- B. End of Cooling**
1. When everything is almost cool, remaining magma will have high silicose content, and quartz will form.
 2. When cooling is complete, everything that cooled at the same time will usually be close to one another (feldspar, mica and quartz cool near one another to make granite).



Earth Science High School Essential Question

PT Brinkman



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Resources in Education ,1992-07 *The Mastery Learning Handbook* Jonathan Bergmann, 2022-10-14 Mastery learning is an instructional approach that empowers every student to progress with confidence Using flexible pacing and targeted supports teachers guide students through a cyclic process of preparation demonstration and formative feedback until there is a mutual agreement between teacher and student that the student is ready to advance In this book educator Jonathan Bergmann a pioneer of the flipped classroom movement walks you step by step through the mastery learning cycle explaining what it entails and providing the templates models and rubrics you need to start using it in your own classroom You ll learn how to Set meaningful measurable and transferable learning objectives that target essential knowledge skills and dispositions Develop a mastery rubric to check for student levels of proficiency Create an assessment plan that ensures positive learning experiences for all Plan and deliver units that incorporate both time shifted direct instruction and collaborative application activities within the classroom space Provide timely differentiated support based on students individual learning needs Informed by trial and error in his own classroom and by discussions with other expert practitioners Bergmann shares commonsense solutions to the major challenges of mastery learning implementation everything from how to manage pacing to how to create multiple versions of tests determine grades and get stakeholder buy in *The Mastery Learning Handbook* is both an introduction to this exciting instructional approach and a practical resource that K 12 teachers can turn to again and again See for yourself just how effective enjoyable and transformative mastery learning can be

Science Curriculum Resource Handbook ,1992

Intersections of Formal and Informal Science Lucy Avraamidou, Wolff-Michael Roth, 2016-03-10 Science learning that takes place between and at the intersections of formal and informal science environments has not been systematically reviewed to offer a comprehensive understanding of the existing knowledge base Bringing together theory and research this volume describes the various ways in which learning science in various settings has been conceptualized as well as empirical evidence to illustrate how science learning in these settings can be supported

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Developing Learning Communities Through Teacher Expertise Giselle O. Martin-Kniep, 2003-10-14 Learn practical methods for developing a collaborative environment where teachers and administrators work together to enhance teachers practices increase student learning and produce valuable school processes

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Proceedings of the High School Conference of ... Horace Adelbert Hollister, 1919

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