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**Rosina Ehmann**



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molecule is one where the charge distribution around the molecule is not symmetric it results from having polar bonds and also a molecular structure where the bond polarities do not cancel

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web the relative electronegativities of the bonded atoms is  $\text{H} < \text{C} < \text{Cl}$  and so the bond moments all point toward the  $\text{Cl}$  end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as  $\text{BF}_3$  trigonal planar  $\text{CH}_4$  tetrahedral  $\text{PF}_5$  trigonal bipyramidal and  $\text{SF}_6$

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