

Chem 8 Practice

1. The balanced equation $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$ tells us that 4.0 mol H_2

- A) reacts with 2.0 mol P_4 .
- B) produces 8.0 mol PH_3 .
- C) cannot react with phosphorus.
- D) produces 2.7 mol PH_3 .
- E) reacts with 4.0 mol P_4 .

2. A 1.0-mol sample of PCl_5 was decomposed according to the equation



How many moles of Cl_2 are formed assuming 100% yield?

- A) 0.5 mol
- B) 1.5 mol
- C) 1.0 mol
- D) 0.7 mol
- E) 2.0 mol

3. The rusting of iron is represented by the equation $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. If you have a 1.04-mol sample of iron, how many moles of Fe_2O_3 will there be when the iron has rusted completely?

- A) 0.527 mol
- B) 0.790 mol
- C) 1.05 mol
- D) 1.54 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.

$$5\text{X}_2(\text{s}) + 2\text{H}_2\text{O}_2(\text{l}) \rightarrow \text{H}_2\text{X}_4(\text{l}) + 2\text{H}_2\text{O}(\text{l})$$

If 175 g of X react completely with hydrogen peroxide to produce 1.43 mol of H_2X_4 , what is the identity of X?

- A) Ge
- B) Mg
- C) Si
- D) Sn
- E) C

5. Refer to the following equation:



How many molecules of NO_2 are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B) 3.64×10^{23}
- C) 1.82×10^{23}
- D) 128
- E) none of these

6. In the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$, how many moles of ammonia would be produced from 1.08 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 3.27 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.543 mol

7. Refer to the following unbalanced equation:

$$\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

What mass of oxygen (O_2) is required to react completely with 13.1 g of C_4H_{10} ?

- A) 2.98×10^2 g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.152 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A) 1.37×10^{23} molecules
- B) 1.03×10^{23} molecules
- C) 4.14×10^{23} molecules
- D) 3.10×10^{23} molecules
- E) 5.17×10^{23} molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane, C_8H_{18} , in the following reaction:



- A) 79.2 g CO_2
- B) 15.8 g CO_2
- C) 88.1 g CO_2
- D) 1.238 g CO_2
- E) 1.400 g CO_2

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction:



- A) 1.03×10^{23} molecules O_2
- B) 2.28×10^{23} molecules O_2
- C) 1.71×10^{23} molecules O_2
- D) 3.42×10^{23} molecules O_2
- E) 5.13×10^{23} molecules O_2

Multiple Choice Stoichiometry Test

Roger Slade, Maureen Slade



Multiple Choice Stoichiometry Test:

This book delves into Multiple Choice Stoichiometry Test. Multiple Choice Stoichiometry Test is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Multiple Choice Stoichiometry Test, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Multiple Choice Stoichiometry Test
- Chapter 2: Essential Elements of Multiple Choice Stoichiometry Test
- Chapter 3: Multiple Choice Stoichiometry Test in Everyday Life
- Chapter 4: Multiple Choice Stoichiometry Test in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Multiple Choice Stoichiometry Test. This chapter will explore what Multiple Choice Stoichiometry Test is, why Multiple Choice Stoichiometry Test is vital, and how to effectively learn about Multiple Choice Stoichiometry Test.
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4. In chapter 3, this book will examine the practical applications of Multiple Choice Stoichiometry Test in daily life. This chapter will showcase real-world examples of how Multiple Choice Stoichiometry Test can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Multiple Choice Stoichiometry Test in specific contexts. This chapter will explore how Multiple Choice Stoichiometry Test is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Multiple Choice Stoichiometry Test. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Multiple Choice Stoichiometry Test.

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web g if the council recommends proceeding with the ordination the church will then set a time for the ordination service u ordination service this may be incorporated into a regular sunday morning or evening service of the church or it may be planned for a special time call to worship hymn of praise oh for a thousand tongues

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