

14.5

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. autoionization |
| | 7. acid solution | 8. pH |

Key Equations

- | | |
|--|------------------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (Eq. 14.1) | 2. $pH = -\log[H_3O^+]$ (Eq. 14.2) |
| 3. $pH = -\log[H_3O^+]$ (Eq. 14.3) | 4. $pH = -\log[H_3O^+]$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete the problem.

1. What product is the _____ of a weak hydroxide ion and a hydronium ion? _____ of the autoionization of water which produces an acidic solution. _____
2. Hydrochloric acid ionizes from _____ to a strong acid. _____
3. _____ autoionization produces the hydronium ion. _____
4. _____ autoionization _____ water. _____
5. _____ autoionization has a very small K_a and _____ autoionization _____
6. _____ autoionization produces a strong acid.
7. _____ autoionization produces a strong base.

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