

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode  $T \gg \theta_M$ ) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute  $\frac{3}{2}R$  to the molar heat capacity. If we denote the number of active rotational modes by  $\nu_R^*$  (so for most molecules at normal temperatures  $\nu_R^* = 2$  for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{3}{2}R + \nu_R^*R + \dots$$

The equipartition value  $C_{V,m} = \frac{3}{2}R + \nu_R^*R + \dots$  is reached when the temperature is high enough so that  $T \gg \theta_M$  for all the modes. However, if the temperature is lower, some of the modes are not active and the heat capacity is lower. The heat capacity is lower because some of the modes are not active and the heat capacity is lower. The heat capacity is lower because some of the modes are not active and the heat capacity is lower. The heat capacity is lower because some of the modes are not active and the heat capacity is lower.

### Solutions to exercises

The heat capacity of a diatomic gas is  $C_{V,m} = \frac{3}{2}R + \nu_R^*R + \nu_V^*R$ . When the temperature is high enough so that  $T \gg \theta_M$  for all the modes, the heat capacity is  $C_{V,m} = \frac{3}{2}R + \nu_R^*R + \nu_V^*R$ .

$$C_{V,m} = \frac{3}{2}R + \nu_R^*R + \nu_V^*R$$

At 298 K,  $\theta_M = 33.5$  K and  $\theta_V = 3300$  K. Since  $T \gg \theta_M$  but  $T \ll \theta_V$ , we have  $\nu_R^* = 2$  and  $\nu_V^* = 0$ .

$$C_{V,m} = \frac{3}{2}R + 2R + 0R = \frac{7}{2}R = 29.1 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 1000 K,  $\theta_M = 33.5$  K and  $\theta_V = 3300$  K. Since  $T \gg \theta_M$  but  $T \ll \theta_V$ , we have  $\nu_R^* = 2$  and  $\nu_V^* = 0$ .

$$C_{V,m} = \frac{3}{2}R + 2R + 0R = \frac{7}{2}R = 29.1 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 3300 K,  $\theta_M = 33.5$  K and  $\theta_V = 3300$  K. Since  $T \gg \theta_M$  and  $T \approx \theta_V$ , we have  $\nu_R^* = 2$  and  $\nu_V^* = 1$ .

$$C_{V,m} = \frac{3}{2}R + 2R + 1R = \frac{9}{2}R = 37.4 \text{ J K}^{-1} \text{ mol}^{-1}$$

# Physical Chemistry 9th Solution Manual

**RC Schank**



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## Unveiling the Energy of Verbal Artistry: An Psychological Sojourn through **Physical Chemistry 9th Solution Manual**

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dec 3 2021 fiches d exercice dictée cm1 à imprimer vous y trouverez ci dessous des exemples de dictée cm1 dictée cm1 conjugaison et les accords la classe de mallory propose une série de courts textes pour cerner les difficultés des élèves en début de l année

*cm1 cm2 dictées flashes chez monsieur paul* - Jun 10 2022

corrigé côté mon ancien restaurer électricité plomberie beaucoup malgré vieux charme certain bien 2 grammaire place les

mots à apprendre dans le tableau des natures tu peux utiliser ton dictionnaire 3 vocabulaire range les

**manuels anciens jean maitron 60 dictées suivies de questions** - May 09 2022

lundi 11 février 2013 jean maitron 60 dictées suivies de questions jean maitron 60 dictées suivies de questions traitées et notées les Éditions ouvrières 1954 entrée en sixième classes de sixième et cinquième certificat d études primaires ouvrage scanné par adelaideaugusta maitron jean