

## 17.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide and hydronium ion concentrations, pH, and pOH, and vice versa, for strong acids and bases.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                                |        |
|-------------------|--------------------------------|--------|
| 1. autoionization | 2. conjugate acid-base pair    | 3. pH  |
| 4. hydronium ion  | 5. hydronium ion concentration | 6. pOH |
|                   | 7. weak acid                   |        |

## Key Equations

- |   |   |
|---|---|
| 1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 17.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 17.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 17.3)    | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 17.4)  |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair represented in the reaction. Also indicate which member of the pair is the stronger acid.

1.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$   
 Identify the conjugate acid-base pair in the autoionization of  $\text{H}_2\text{SO}_4$ . Which species is present in higher concentration,  $\text{H}_2\text{SO}_4$  or  $\text{HSO}_4^-$ ?
2.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$   
 Identify the conjugate acid-base pair in the autoionization of  $\text{H}_2\text{SO}_4$ . Which species is present in higher concentration,  $\text{H}_2\text{SO}_4$  or  $\text{HSO}_4^-$ ?
3.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$   
 Identify the conjugate acid-base pair in the autoionization of  $\text{H}_2\text{SO}_4$ . Which species is present in higher concentration,  $\text{H}_2\text{SO}_4$  or  $\text{HSO}_4^-$ ?
4.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$   
 Identify the conjugate acid-base pair in the autoionization of  $\text{H}_2\text{SO}_4$ . Which species is present in higher concentration,  $\text{H}_2\text{SO}_4$  or  $\text{HSO}_4^-$ ?
5.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$   
 Identify the conjugate acid-base pair in the autoionization of  $\text{H}_2\text{SO}_4$ . Which species is present in higher concentration,  $\text{H}_2\text{SO}_4$  or  $\text{HSO}_4^-$ ?

6.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$
7.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$
8.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$
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10.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$
11.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$
12.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$
13.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$
14.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$

# 19 2 Hydrogen Ions And Acidity Answer Key

**L Reisser**



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