

Unit 8 Review

1. The balanced equation $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$ tells us that 4.0 mol H_2

A) reacts with 2.0 mol P_4 .
 B) produces 8.0 mol PH_3 .
 C) cannot react with phosphorus.
 D) produces 2.7 mol PH_3 .
 E) reacts with 4.0 mol P_4 .

2. A 1.5-mol sample of Fe_2O_3 was decomposed according to the equation:

$$2\text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}_2\text{O}(\text{s}) + 3\text{O}_2(\text{g})$$

How many moles of O_2 are formed assuming 100% yield?

A) 0.9 mol
 B) 1.5 mol
 C) 1.2 mol
 D) 0.7 mol
 E) 2.0 mol

3. The rusting of iron is represented by the equation $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. If you have a 1.04-mol sample of iron, how many moles of Fe_2O_3 will there be after the iron has rusted completely?

- A) 0.527 mol
 B) 0.790 mol
 C) 1.05 mol
 D) 1.58 mol
 E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element:

$$8\text{X}_2(\text{s}) + 2\text{B}_2\text{O}_3(\text{s}) \rightarrow \text{B}_2\text{X}_4(\text{s}) + 3\text{X}_2\text{O}_3(\text{g})$$

If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of B_2X_4 , what is the identity of X?

- A) Ge
 B) Mg
 C) Si
 D) Sn
 E) C

5. Refer to the following equation:

$$4\text{NH}_3(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$$

How many molecules of NO are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
 B) 3.84×10^{23}
 C) 1.82×10^{24}
 D) 1.39
 E) none of these

6. In the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$, how many moles of ammonia would be produced from 1.09 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
 B) 3.27 mol
 C) 0.727 mol
 D) 2.18 mol
 E) 0.363 mol

7. Refer to the following unbalanced equation:

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

What mass of oxygen (O_2) is required to react completely with 13.1 g of C_2H_6 ?

- A) 2.99×10^2 g
 B) 23.1 g
 C) 4.86 g
 D) 49.2 g
 E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A) 1.37×10^{23} molecules
 B) 1.03×10^{23} molecules
 C) 4.14×10^{23} molecules
 D) 3.10×10^{23} molecules
 E) 5.17×10^{23} molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane, C_8H_{18} , in the following reaction:



- A) 79.2 g CO_2
 B) 158 g CO_2
 C) 89.1 g CO_2
 D) 1.258 g CO_2
 E) 1.400 g CO_2

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction:



- A) 1.03×10^{23} molecules O_2
 B) 2.28×10^{23} molecules O_2
 C) 1.71×10^{23} molecules O_2
 D) 3.42×10^{23} molecules O_2
 E) 5.13×10^{23} molecules O_2

Multiple Choice Stoichiometry Test

Paul Foglino, Princeton Review (Firm)

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