

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH, and vice versa, for strong acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. strong acid/base |
| 4. hydronium ion | 5. weak acid/base | 6. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (17.1) | 2. $\text{pH} = -\log[\text{H}^+]$ (17.2) |
| 3. $\text{pH} + \text{pOH} = 14$ (17.3) | 4. $\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$ (17.5) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete the problem.

1. What product is the _____ of a weak hydroxide ion and a hydronium ion? _____ of the conjugate acid-base pair.
2. _____ of a strong acid-base pair is _____.
3. _____ of a weak acid-base pair is _____.
4. _____ of a strong acid-base pair is _____.
5. _____ of a weak acid-base pair is _____.
6. _____ of a strong acid-base pair is _____.
7. _____ of a weak acid-base pair is _____.
8. _____ of a strong acid-base pair is _____.

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