

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from K_w and K_a or K_b and the concentration of a weak acid or base.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. K_a | 6. K_b |
| | 7. weak base | |

Key Equations

- | | |
|--|---|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ | 4. $\text{pOH} = -\log[\text{OH}^-]$ |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions. Show calculations for all problems that require calculations.

1. What products are formed when hydrochloric acid and sodium hydroxide react? Write the balanced chemical equation for this reaction and identify the products as strong electrolyte, weak electrolyte, or none.
2. Hydrochloric acid has a molar concentration of 0.010 M. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of this solution.
3. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of acetic acid.
4. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of sodium acetate.
5. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of sodium hydroxide.
6. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of hydrochloric acid.
7. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of sodium hydroxide.

8. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of acetic acid.
9. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of sodium acetate.
10. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of sodium hydroxide.
11. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of hydrochloric acid.
12. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of sodium hydroxide.
13. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of hydrochloric acid.
14. Calculate the $[\text{H}_3\text{O}^+]$, $[\text{OH}^-]$, and pH of a 0.010 M solution of sodium hydroxide.

19 2 Hydrogen Ions And Acidity Section Review Answers

Emilie Sanchez



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Table of Contents 19 2 Hydrogen Ions And Acidity Section Review Answers

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Section Review Answers
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Section Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Personalized Recommendations

- 19 2 Hydrogen Ions And Acidity Section Review Answers User Reviews and Ratings
- 19 2 Hydrogen Ions And Acidity Section Review Answers and Bestseller Lists
- 5. Accessing 19 2 Hydrogen Ions And Acidity Section Review Answers Free and Paid eBooks
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Section Review Answers eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Budget-Friendly Options
- 6. Navigating 19 2 Hydrogen Ions And Acidity Section Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Section Review Answers
- 8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Section Review Answers
- 9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Section Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Section Review Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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