

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide and hydronium ion concentrations and pH and pOH and the associated logarithmic functions.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                               |                   |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair   | 3. acid solution  |
| 4. autoionization | 5. hydronium ion ( $H_3O^+$ ) | 6. autoionization |
|                   | 7. acid solution              | 8. pH             |

## Key Equations

- |   |                         |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$                             | 4. $pOH = -\log[OH^-]$  |

## Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate acid/base.

1. Water autoionizes to  $H_3O^+$  and  $OH^-$ . Hydroxide ion acts as the conjugate base of the autoionization of water which produces an acidic solution. \_\_\_\_\_
2. Hydrochloric acid dissociates into  $H^+$  and  $Cl^-$ . Chloride ion is the conjugate base of the acid. The resulting solution is acidic. \_\_\_\_\_
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1. conjugate acid
2. conjugate base
3. conjugate acid
4. conjugate base
5. conjugate acid
6. conjugate base
7. conjugate acid
8. conjugate base
9. conjugate acid
10. conjugate base
11. conjugate acid
12. conjugate base
13. conjugate acid
14. conjugate base
15. conjugate acid
16. conjugate base
17. conjugate acid
18. conjugate base

# 19 2 Hydrogen Ions And Acidity Answer Key

**Christian G. Meyer**



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