

Chemistry: 2nd Semester Exam

Select the correct oxidation numbers for the following elements.

- Oxygen in O_2
 - 2
 - +2
 - 1
 - +1
 - 0
- Sodium in $NaCl$
 - 2
 - +2
 - 1
 - +1
 - 0
- Phosphorus in H_3PO_4
 - 5
 - 3
 - 0
 - +3
 - +5
- H in H_2O
 - 2
 - +2
 - 1
 - +1
 - 0
- O in H_2O
 - 2
 - +2
 - 1
 - +1
 - 0
- In which substance does bromine have an oxidation number of +1?
 - Br_2
 - HBr
 - $HBrO$
 - $HBrO_2$
- Which species is the oxidizing agent in the following reaction? $Cl_2(aq) + 2I^-(aq) \rightarrow I_2(aq) + 2Cl^-(aq)$
 - Cl_2
 - $2I^-$
 - I_2
 - $2Cl^-$
- Consider the following: $2MnO_4^- + 5H_2SO_3 \rightarrow 2Mn^{2+} + 3H_2O + 5SO_4^{2-} + 4H^+$
Which of the following elements does not experience a change in oxidation number?
 - S in H_2SO_3
 - O in MnO_4^-
 - Mn in MnO_4^-
- In which of the following unbalanced equations does chromium undergo oxidation?
 - $Cr^{3+} \rightarrow Cr$
 - $Cr^{2+} \rightarrow Cr^{3+}$
 - $Cr^{3+} \rightarrow Cr_2O_7^{2-}$
 - $CrO_4^{2-} \rightarrow Cr_2O_7^{2-}$
- In a redox reaction, the species which loses electrons is _____.
 - reduced
 - oxidized
 - neutralized
- Consider the following: $Cl_2 \rightarrow ClO_2^-$
The balanced half reaction is:
 - $Cl_2 + 3H_2O \rightarrow 2ClO_2^- + 6H^+ + 5e^-$
 - $Cl_2 + 3H_2O \rightarrow 2ClO_2^- + 6H^+ + 4e^-$
 - $Cl_2 + 6H_2O + 2e^- \rightarrow 2ClO_2^- + 6H_2$
 - $Cl_2 + 6H_2O \rightarrow 2ClO_2^- + 12H^+ + 10e^-$

Label the Acid, Base, Conjugate Acid, and Conjugate Base



- $HC_2H_3O_2$ serves as the
 - Acid
 - Base
 - Conjugate Acid
 - Conjugate Base
- H_2O serves as the
 - Acid
 - Base
 - Conjugate Acid
 - Conjugate Base
- H_3O^+ serves as the
 - Acid
 - Base
 - Conjugate Acid
 - Conjugate Base
- $C_2H_3O_2^-$ serves as the
 - Acid
 - Base
 - Conjugate Acid
 - Conjugate Base

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