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web step 4 the reactant that produces a smaller amount of product is the limiting reagent  $\text{Mg}$   $\text{Mg}$  produces less  $\text{MgO}$   $\text{MgO}$  than does  $\text{O}_2$   $\text{O}_2$   $3.98\text{ g MgO}$  vs  $25.2\text{ g MgO}$  therefore  $\text{Mg}$  is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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