

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydroxide ion | 5. hydronium ion | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{pH} = -\log[\text{H}^+]$ (eq. 14.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (eq. 14.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (eq. 14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the acid-base reaction given in the following examples.

1. What product is the _____ of the hydroxide ion and the hydronium ion? _____
2. The conjugate base of the hydronium ion is the _____ of the hydronium ion. _____
3. _____ is the conjugate acid of the hydroxide ion. _____
4. _____ is the conjugate base of the hydronium ion. _____
5. _____ is the conjugate acid of the hydroxide ion. _____
6. _____ is the conjugate base of the hydronium ion. _____
7. _____ is the conjugate acid of the hydroxide ion. _____
8. _____ is the conjugate base of the hydronium ion. _____
9. _____ is the conjugate acid of the hydroxide ion. _____
10. _____ is the conjugate base of the hydronium ion. _____

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

19 2 Hydrogen Ions And Acidity Answer Key

R Barnett



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