

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH⁻] and [H⁺].
3. Calculate the pH of weak (or strong) acids.

Vocabulary

- | | | |
|-------------------|---|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. hydronium ion (H ₃ O ⁺) | 6. autoionization |
| | 7. acid solution | 8. pH |

Key Equations

- | | |
|---|---------------------------------|
| 1. $[H^+][OH^-] = 1.0 \times 10^{-14}$ (Eq. 14.1) | 2. $pH = -\log[H^+]$ (Eq. 14.2) |
| 3. $pH = -\log[H^+]$ (Eq. 14.3) | 4. $pH = -\log[H^+]$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in (a). (b) is a special case of (a) and is not included in the list.

- Water autoionizes to H_3O^+ and OH^- in pure water. Identify the autoionization reaction and the conjugate acids of these species produced in this reaction. (b) H_2O and H_3O^+ .
- Phosphoric acid dissociates to $H_2PO_4^-$ and H^+ in water. Identify $H_2PO_4^-$ as a conjugate acid or base of the water species. (b) $H_2PO_4^-$ is a conjugate base.
- H_2O is a conjugate base of H_3O^+ and a conjugate acid of OH^- .
- H_2O is a conjugate base of H_3O^+ and a conjugate acid of OH^- .
- H_2O is a conjugate base of H_3O^+ and a conjugate acid of OH^- .
- H_2O is a conjugate base of H_3O^+ and a conjugate acid of OH^- .
- H_2O is a conjugate base of H_3O^+ and a conjugate acid of OH^- .

- conjugate acid
- conjugate base

19 2 Section Review Hydrogen Ions And Acidity

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