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molecule is one where the charge distribution around the molecule is not symmetric it results from having polar bonds and also a molecular structure where the bond polarities do not cancel

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web when the atoms in a bond are not the same the bonding electrons are shared and the bond is electronegativities equally unequally non polar polar unequally polar the degree of polarity of a bond between any two atoms is determined by consulting a table of

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web chemical bonds hold molecules together and create temporary connections that are essential to life types of chemical bonds including covalent ionic and hydrogen bonds and london dispersion forces introduction living things are made up of atoms but in most cases those atoms aren't just floating around individually

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web the relative electronegativities of the bonded atoms is h c cl and so the bond moments all point toward the cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as BF_3 trigonal planar CH_4 tetrahedral PF_5 trigonal bipyramidal and SF_6

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