

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode  $T \gg \theta_M$ ) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute  $\frac{3}{2}R$  to the molar heat capacity. If we denote the number of active rotational modes by  $\nu_R^\ddagger$  (so for most molecules at normal temperatures  $\nu_R^\ddagger = 2$  for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$$

The equipartition value  $C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R$  is reached when the temperature is high enough so that  $T \gg \theta_M$  for all the modes. However, if the temperature is lower, some of the modes are not fully active and the heat capacity is lower than the equipartition value. The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

### Solutions to exercises

**16.1** The vibrational energy is given by  $\epsilon = \frac{h\nu}{2} + \frac{h\nu}{2} + \dots$ . The vibrational energy is  $\epsilon = \frac{h\nu}{2} + \frac{h\nu}{2} + \dots$ .

$$\epsilon = \frac{h\nu}{2} + \frac{h\nu}{2} + \dots$$

**16.2** The vibrational energy is given by  $\epsilon = \frac{h\nu}{2} + \frac{h\nu}{2} + \dots$ .

$$\epsilon = \frac{h\nu}{2} + \frac{h\nu}{2} + \dots$$

**16.3** The vibrational energy is given by  $\epsilon = \frac{h\nu}{2} + \frac{h\nu}{2} + \dots$ .

$$\epsilon = \frac{h\nu}{2} + \frac{h\nu}{2} + \dots$$

**16.4** The vibrational energy is given by  $\epsilon = \frac{h\nu}{2} + \frac{h\nu}{2} + \dots$ .

**16.5** The vibrational energy is given by  $\epsilon = \frac{h\nu}{2} + \frac{h\nu}{2} + \dots$ .

# Physical Chemistry Atkins 9th Solutions Manual

**Puja Mehta**



**Physical Chemistry Atkins 9th Solutions Manual:**

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In a digitally-driven world where displays reign supreme and instant interaction drowns out the subtleties of language, the profound techniques and emotional subtleties hidden within words often get unheard. However, located within the pages of **Physical Chemistry Atkins 9th Solutions Manual** a captivating fictional treasure sporting with natural thoughts, lies a fantastic quest waiting to be undertaken. Composed by a talented wordsmith, that wonderful opus invites viewers on an introspective trip, gently unraveling the veiled truths and profound impact resonating within the very fabric of each and every word. Within the emotional depths with this poignant review, we will embark upon a genuine exploration of the book is key themes, dissect their captivating writing model, and succumb to the powerful resonance it evokes strong within the recesses of readers hearts.

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