

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide and hydronium ion concentrations from K_b and K_a values and the concentration of a weak acid or base.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydroxide ion | 5. hydronium ion (H_3O^+) | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|--|
| 1. $K_a = \frac{[H_3O^+][A^-]}{[HA]}$ (Eq. 17.1) | 2. $K_b = \frac{[OH^-][BH^+]}{[B]}$ (Eq. 17.2) |
| 3. $pH = -\log[H_3O^+]$ (Eq. 17.3) | 4. $pOH = -\log[OH^-]$ (Eq. 17.4) |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that are based on the following acid dissociation constants (see Appendix C):

1. What products are **1** H_2O and H_3O^+ formed in the autoionization of H_2O ? The autoionization of H_2O is represented by the equation $H_2O(l) \rightleftharpoons H_2O(l) + H_2O(l)$.
2. Hydrofluoric acid dissociates from **2** H_2O to H_3O^+ and F^- . Identify **3** H_2O as the conjugate base of the acid and identify **4** H_3O^+ as the conjugate acid of F^- . What is the K_a of H_2O ?
3. **5** H_2O is the conjugate base of H_3O^+ and OH^- is the conjugate acid of H_2O . What is the K_b of H_2O ?
4. **6** H_2O is the conjugate base of H_3O^+ and OH^- is the conjugate acid of H_2O . What is the K_a of H_2O ?
5. **7** H_2O is the conjugate base of H_3O^+ and OH^- is the conjugate acid of H_2O .

8. **8** H_2O is the conjugate base of H_3O^+ .
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19 2 Hydrogen Ions And Acidity Answer Key

S Ashworth



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