
OSHA 10 - Introduction to OSHA

Questions and Answers 100% Correct

Question 1: Is Adrian Smith, one of 3 employees of ABC Landscaping, covered by OSHA? - ANSWER-A. Yes

Question 2: Is Taylor Dell, an accountant in business for herself, covered by OSHA? - ANSWER-B. No

Question 3: Is Rob Jones, one of 10 carpenters working for Woody, Inc., covered by OSHA? - ANSWER-A. Yes

Question 4: Why was OSHA necessary?

- A. Workplace injuries, illnesses and death were increasing
- B. No uniform or Comprehensive law existed to protect against workplace hazards
- C. None of the Above
- D. Both A and B - ANSWER-D. Both A and B

Question 5: What is OSHA's mission?

- A. To save lives
- B. To prevent injuries
- C. Protect the health of America's workers
- D. All the answers are correct - ANSWER-D. All the answers are correct

Question 6: Why is this training important?

- A. You'll know your rights if there are hazards in your workplace.
- B. You'll know where to get help.
- C. You'll know if your employer is complying with OSHA standards.
- D. All of the answers are correct - ANSWER-D. All of the answers are correct

Question 1: What are some of the things you might find on an SDS?

- A. Chemical Name
- B. You'll know where to get help
- C. Fire and Explosion hazards
- D. Precautions for safe use and handling
- E. All of the Above - ANSWER-E. All of the Above

Question 2: What are some worker rights related to injury and illness reporting?

- A. Know how to report injuries/illnesses, see the annual summary and review the log
- B. Right to file a complaint
- C. Right to participate in an OSHA inspection

Osha Questions And Answers Guide

Jin-Ying Zhang



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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction

with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens

very rarely because bromine is an ortho para director

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno₃ h₂so₄ in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no₂⁺ which acts as the electrophilic species

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