

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs in these pairs of hydrolyzed acid or base species, respectively.
2. Calculate hydrolysis and neutralization constants (K_b and K_a) for weak acids and bases, respectively, given K_w .
3. Calculate the pH of weak acid or base solutions.

Vocabulary

- | | | |
|-------------------|---|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. hydrolysis | 5. hydrolysis constant (K_b or K_a) | 6. neutralization |
| | 7. weak base/acid | 8. pH |

Key Equations

- | | |
|---|-------------------------------------|
| 1. $K_w = [H_3O^+][OH^-] = 1.0 \times 10^{-14}$ | 2. $K_b = \frac{[OH^-][BH^+]}{[B]}$ |
| 3. $pH = -\log[H_3O^+]$ | 4. $pOH = -\log[OH^-]$ |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that are based on the following data: $K_w = 1.0 \times 10^{-14}$ at 25°C.

1. What products are _____ when hydrochloric acid (HCl) is neutralized with _____? The conjugate base of this acid is _____.
2. _____ is the conjugate base of _____.
3. _____ is the conjugate acid of _____.
4. _____ is the conjugate base of _____.
5. _____ is the conjugate acid of _____.
6. _____ is the conjugate acid of _____.
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17. _____ is the conjugate base of _____.
18. _____ is the conjugate acid of _____.
19. _____ is the conjugate base of _____.
20. _____ is the conjugate acid of _____.

19 2 Hydrogen Ions And Acidity Answer Key

Cecie Starr, John Alcock



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