

## Chemistry 2<sup>nd</sup> Semester Final Review **KEY**

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

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Gas Laws (Ch. 13):

*Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)*

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)  
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$   
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$   
Since all units are standard (L, atm), we do not need to convert the units before proceeding.  
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
T goes away (not mentioned in problem)  
 $P_1V_1=P_2V_2$   
Isolate for P<sub>2</sub>:  $P_2=(P_1V_1)/V_2$   
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$  (L cancelled and it's 2 sig figs)
2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)  
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$   
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$   
We need to convert from Celsius to Kelvins before proceeding.  
 $T_1=34.75 + 273=307.75 \text{ K}$   
 $T_2=78.41 + 273=351.41 \text{ K}$   
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
Get rid of P  
 $V_1/T_1=V_2/T_2$   
Isolate V<sub>2</sub>:  $V_2=(V_1/T_1)T_2$   
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?  
The combined gas law won't work here. You only have one condition and nothing changes.  
 $PV=nRT$  will work. (n= moles and R= 0.08206 L atm/K mol)  
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$  OOPS we need to convert from grams to moles!,  $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$   
n: Hydrogen gas is diatomic!  
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$   
Isolate V:  $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
4. What volume is occupied by 56.75 g of oxygen gas at STP?

# 2nd Semester Chemistry Final Exam Study Guide

**N Noddings**



## **2nd Semester Chemistry Final Exam Study Guide:**

## Reviewing **2nd Semester Chemistry Final Exam Study Guide**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**2nd Semester Chemistry Final Exam Study Guide**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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