

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode  $T \gg \theta_M$ ) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute  $\frac{3}{2}R$  to the molar heat capacity. If we denote the number of active rotational modes by  $\nu_R^\ddagger$  (so for most molecules at normal temperatures  $\nu_R^\ddagger = 2$  for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{1}{2}R(3 + \nu_R^\ddagger + \nu_V^\ddagger)$$

The equipartition value of  $C_{V,m}$  is  $\frac{1}{2}R(3 + \nu_R^\ddagger + \nu_V^\ddagger)$  and is the value that would be obtained if all the modes were active. However, if the temperature is not high enough to excite all the modes, the heat capacity will be less than the equipartition value. The heat capacity will approach the equipartition value as the temperature increases. The heat capacity will approach the equipartition value as the temperature increases.

### Solutions to exercises

The heat capacity of a diatomic gas is  $C_{V,m} = \frac{1}{2}R(3 + \nu_R^\ddagger + \nu_V^\ddagger)$ . The equipartition value is  $\frac{1}{2}R(3 + \nu_R^\ddagger + \nu_V^\ddagger)$ .

$$C_{V,m} = \frac{1}{2}R(3 + 2 + \nu_V^\ddagger)$$

At 298 K,  $\theta_V = 3300$  K, so  $\nu_V^\ddagger = 0$ .

$$C_{V,m} = \frac{1}{2}R(3 + 2 + 0) = \frac{5}{2}R = 20.8 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 1000 K,  $\theta_V = 3300$  K, so  $\nu_V^\ddagger = 0$ .

$$C_{V,m} = \frac{1}{2}R(3 + 2 + \frac{1}{2}) = \frac{7}{4}R = 29.1 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 2000 K,  $\theta_V = 3300$  K, so  $\nu_V^\ddagger = 0$ .

$$C_{V,m} = \frac{1}{2}R(3 + 2 + 0) = \frac{5}{2}R = 20.8 \text{ J K}^{-1} \text{ mol}^{-1}$$

# Physical Chemistry 9th Solution Manual

**Chao Zhang**



**Physical Chemistry 9th Solution Manual:**

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In a world taken by monitors and the ceaseless chatter of fast transmission, the melodic elegance and psychological symphony created by the written word frequently disappear in to the background, eclipsed by the relentless noise and distractions that permeate our lives. But, set within the pages of **Physical Chemistry 9th Solution Manual** a marvelous literary value filled with raw emotions, lies an immersive symphony waiting to be embraced. Crafted by an outstanding musician of language, this captivating masterpiece conducts visitors on a mental trip, skillfully unraveling the hidden melodies and profound impact resonating within each cautiously constructed phrase. Within the depths of this moving examination, we shall examine the book is key harmonies, analyze their enthralling publishing style, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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