1. The balanced equation $P_4(s) + 6H_2(g) \rightarrow 4PH_3(g)$ tells us that 4.0 mol H_2
   A) reacts with 2.0 mol P_4.
   B) produces 6.0 mol PH_3.
   C) cannot react with phosphorus.
   D) produces 2.7 mol PH_3.
   E) reacts with 4.0 mol P_4.

3. A 1.3-mol sample of $KClO_3$ was decomposed according to the equation
   $2KClO_3(s) \rightarrow 2KCl(s) + 3O_2(g)$
   How many moles of O_2 are formed assuming 100% yield?
   A) 0.9 mol
   B) 1.1 mol
   C) 1.3 mol
   D) 0.7 mol
   E) 2.0 mol

5. The rusting of iron is represented by the equation $4Fe + 3O_2 \rightarrow 2Fe_2O_3$. If you have a 1.700-mol sample of iron, how many moles of Fe_2O_3 will there be after the iron has rusted completely?
   A) 0.527 mol
   B) 0.790 mol
   C) 1.05 mol
   D) 1.58 mol
   E) 2.37 mol

7. Consider the following reaction, where X represents an unknown element.
   $6HX(s) + 2B_2O_3(s) \rightarrow B_4X_6(s) + 3O_2(g)$
   If 1.75 g of X react to completely with 0.4 moles of oxygen to produce 2.47 mol of B_4X_6, what is the identity of X?
   A) Ge
   B) Mg
   C) Si
   D) Al
   E) C

9. Refer to the following equation:
   $4NH_3(g) + 7O_2(g) \rightarrow 4NO_2(g) + 6H_2O(g)$
   How many molecules of NO_2 are produced when 3.02 mol of ammonia is completely reacted?
   A) 12.08
   B) 3.64×10^{24}
   C) 1.82×10^{24}
   D) 139
   E) none of these

11. In the reaction $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$, how many moles of ammonia would be produced from 1.09 mol of hydrogen and excess nitrogen?
   A) 1.62 mol
   B) 3.27 mol
   C) 0.727 mol
   D) 2.18 mol
   E) 0.363 mol

13. Refer to the following **unbalanced** equation: $C_3H_{8(g)} + O_{2(g)} \rightarrow CO_{2(g)} + H_2O$
   What mass of oxygen (O_2) is required to react completely with 13.1 g of C_3H_8?
   A) 3.98×10^2 g
   B) 33.1 g
   C) 4.96 g
   D) 66.2 g
   E) 0.132 g

15. How many molecules of carbon dioxide will be formed if 2.37 g of propane is burned in the following reaction?
   $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(g)$
   A) 1.37×10^{23} molecules
   B) 1.03×10^{23} molecules
   C) 4.14×10^{23} molecules
   D) 3.10×10^{23} molecules
   E) 9.17×10^{23} molecules

17. Calculate the mass of carbon dioxide produced from 23.7 g of octane, C_8H_{18}, in the following reaction.
   $2C_8H_{18}(g) + 25O_2(g) \rightarrow 16CO_2(g) + 18H_2O(g)$
   A) 78.3 g CO_2
   B) 156 g CO_2
   C) 69.1 g CO_2
   D) 1.23E3 g CO_2
   E) 1.80E3 g CO_2

19. Calculate the molecules of oxygen required to react with 16.2 g of rubber in the following reaction.
   $2S(s) + 3O_2(g) \rightarrow 2SO_3(g)$
   A) 1.03×10^{24} molecules O_2
   B) 2.08×10^{24} molecules O_2
   C) 1.71×10^{24} molecules O_2
   D) 3.42×10^{24} molecules O_2
   E) 5.13×10^{24} molecules O_2

Multiple Choice Stoichiometry Test

David Wilson

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