

Introduction to Chemical Reactions

A chemical reaction is a process in which substances undergo chemical change. In a chemical change, the atoms rearrange and new substances are formed. The new substances will have new and different properties. Some clues that a chemical reaction has occurred are the production of a gas, a precipitate (solid) forms, or there is a change in color. In an exothermic reaction, like when a candle burns, heat is given off. In an endothermic reaction, like when vinegar is added to baking soda, heat is lost and the substance feels colder. In a chemical reaction, the substances that undergo the chemical changes are called reactants. The substances that form are called the products. This reaction can be shown in a chemical equation. The regular-sized numbers in front of a chemical formula are called coefficients. Coefficients tell you how many molecules of that substance you have. Coefficients apply to every atom they are in front of. The little numbers that you see after some elements are called subscripts. Subscripts apply only to the atom that they come after. The arrow in the equation is often read as "yields". One thing to remember about chemical equations is that they must always be balanced due to the Law of Conservation of Mass. Since matter is neither created nor destroyed during any physical or chemical change, you must always have the same number of each type of atom on both sides of the equation. To balance a chemical equation, you are allowed to adjust the coefficients. You are NEVER allowed to change the subscripts!



Label It

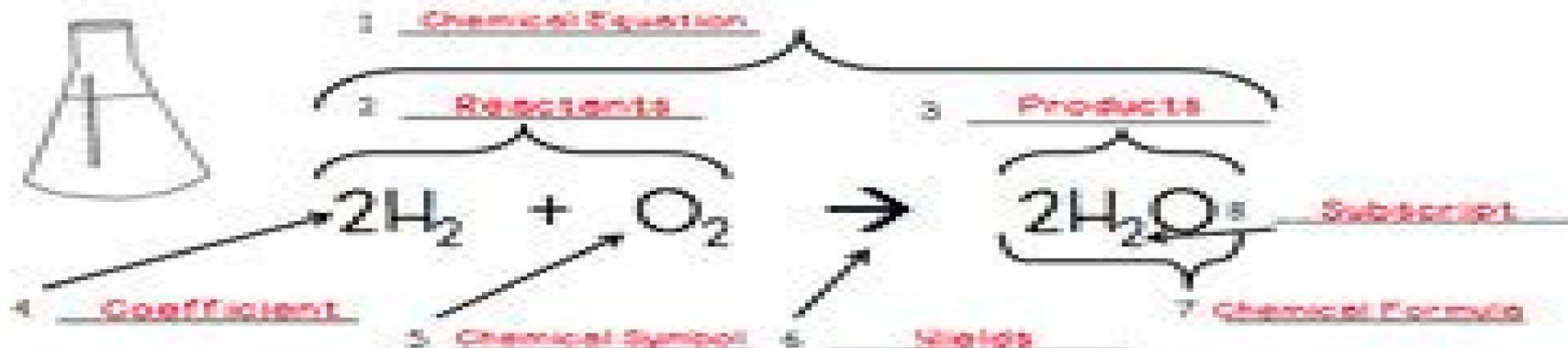
Label the diagram below using the words from the word bank.

Chemical Equation
Products

Chemical Formula
Reactants

Chemical Symbol
Subscript

Coefficient
Yield



An Introduction To Chemical Science

Lei Shi



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