

# **OCTOBER NOVEMBER 2013, D MATH 4024, 12**

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 $\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$   
 $\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$   
 $\text{Pb} + \text{OH}^- \rightarrow 2\text{HCl} + \text{H}_2\text{O} + \text{PbCl}_2$   
 $\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{CH}_4 + 2\text{O}_2$   
 $\text{CO}_2 + 2\text{H}_2\text{O} \rightarrow \text{C}_3\text{H}_8 + 5\text{O}_2$   
 $3\text{CO}_2 + 4\text{H}_2\text{O} \rightarrow \text{C}_8\text{H}_{18} + 25\text{O}_2$   
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 $2\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{CH}_4 + 2\text{O}_2$   
 $\text{CO}_2 + 2\text{H}_2\text{O} \rightarrow \text{C}_3\text{H}_8 + 5\text{O}_2$   
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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$   
 $\text{HNO}_3 + \text{H}_2\text{O} \rightarrow 2\text{H}_2\text{O} + \text{O}_2 + 3\text{H}$

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products and direction of the reaction balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow

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