

DRAFT RELEASED FRQ ANSWERS AP CHEMISTRY 2014*(numbers in parentheses and red, represent guesstimate points)***Question 1**

(ii) Potassium and nitrate ions are in solution at the beginning of the reaction and at the end, do not change in any way (i.e., they are spectators), and as such can be omitted from the equation. (1)

(b) In order to ensure that the mass of the precipitate includes only the mass of the solid and none of the water, repeated drying will ensure that all of the water is removed before a final mass is recorded. (1)

(c) Less than. If the lead(II) nitrate were added in the stoichiometric ratio, the conc. of K^{+} and NO_3^{-} would be 1:1, but since the lead(II) nitrate is added in excess, the nitrate ions are present in the larger conc. (1)(d) $1.698 - 1.462 = 0.236$ g of precipitate.

$$\text{Moles of precipitate} = \frac{0.236 \text{ g}}{(207.2 + 126.91 + 126.91 \text{ g mol}^{-1})} \frac{8.41 \times 10^{-2}}{207.2 + 126.91 + 126.91 \text{ g mol}^{-1}} = 5.12 \times 10^{-4} \text{ moles} \quad (1)$$

(e) Mass of I^{-} = (2) (moles of PbI_2) (126.91 g mol⁻¹) = 0.130 g

$$\text{Mass \% of I}^{-} = \frac{0.130 \text{ g}}{0.425 \text{ g}} \times 100 \frac{0.130}{0.425} (100) = 30.6\% \quad (2)$$

(f) The same since the water plays no part in the mass of I^{-} present (this assumes all of the I^{-} is still precipitated by excess lead (II) nitrate, the tablets are the same and that the precipitate is dried completely as before etc.). (1)(g) (i) Yes, since like lead(II) ions, silver ions also form an insoluble precipitate with iodide ions and since the K_{sp} is sufficiently small to suggest a precipitation will occur. (1)

(ii) No, IF after subtracting the mass of the dried precipitate + filter paper, from the mass of filter paper, yields a mass that has less than three significant figures, or

Yes, IF after subtracting the mass of the dried precipitate + filter paper, from the mass of filter paper, yields a mass that has at least three significant figures. (1)

(No is the expected answer I feel, BUT it does depend on the mass of the tablet IMO).

2014 Released Form Chemistry

Judd E. Hollander



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