

Unit 8 Review

1. The balanced equation $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$ tells us that 4.0 mol H_2

- A) reacts with 2.0 mol P_4 .
- B) produces 8.0 mol PH_3 .
- C) cannot react with phosphorus.
- D) produces 2.7 mol PH_3 .
- E) reacts with 4.0 mol P_4 .

2. A 1.0-mol sample of Fe_2O_3 was decomposed according to the equation:



How many moles of O_2 are formed assuming 100% yield?

- A) 0.9 mol
- B) 1.5 mol
- C) 1.2 mol
- D) 0.7 mol
- E) 2.0 mol

3. The rusting of iron is represented by the equation $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. If you have a 1.00-mol sample of iron, how many moles of Fe_2O_3 will there be after the iron has rusted completely?

- A) 0.527 mol
- B) 0.750 mol
- C) 1.00 mol
- D) 1.50 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.



If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of B_2X_6 , what is the identity of X?

- A) Ca
- B) Mg
- C) Si
- D) Sn
- E) C

5. Refer to the following equation:



How many molecules of NO are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B) 3.84×10^{23}
- C) 1.82×10^{23}
- D) 1.39
- E) none of these

6. In the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow$

$2\text{NH}_3(\text{g})$, how many moles of ammonia would be produced from 1.00 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 3.37 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.343 mol

7. Refer to the following unbalanced



What mass of oxygen (O_2) is required to react completely with 13.1 g of C_2H_6 ?

- A) 2.98×10^2 g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A) 1.37×10^{23} molecules
- B) 1.00×10^{23} molecules
- C) 4.14×10^{23} molecules
- D) 3.10×10^{23} molecules
- E) 5.17×10^{23} molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane, C_8H_{18} , in the following reaction.



- A) 79.2 g CO_2
- B) 158 g CO_2
- C) 89.1 g CO_2
- D) 1.258 g CO_2
- E) 1.400 g CO_2

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction.



- A) 1.03×10^{23} molecules O_2
- B) 2.28×10^{23} molecules O_2
- C) 1.71×10^{23} molecules O_2
- D) 3.42×10^{23} molecules O_2
- E) 5.13×10^{23} molecules O_2

Multiple Choice Stoichiometry Test

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Multiple Choice Stoichiometry Test:

Cambridge Checkpoints VCE Chemistry Units 3 and 4 2013 Roger Slade, Maureen Slade, 2012-12-14 Cambridge Checkpoints VCE are updated regularly to provide you with the most up to date exam preparation available *Cracking the AP Chemistry Exam, 2012 Edition* Paul Foglino, 2011-08-02 Provides techniques for achieving high scores on the AP chemistry exam and includes two full length practice tests

Roadmap to the Regents Sasha Alcott, 2003 If Students Need to Know It It's in This Book This book develops the chemistry skills of high school students It builds skills that will help them succeed in school and on the New York Regents Exams Why The Princeton Review We have more than twenty years of experience helping students master the skills needed to excel on standardized tests Each year we help more than 2 million students score higher and earn better grades We Know the New York Regents Exams Our experts at The Princeton Review have analyzed the New York Regents Exams and this book provides the most up to date thoroughly researched practice possible We break down the test into individual skills to familiarize students with the test's structure while increasing their overall skill level We Get Results We know what it takes to succeed in the classroom and on tests This book includes strategies that are proven to improve student performance We provide a breakdown of the skills based on New York standards and objectives hundreds of practice questions organized by skill two complete practice New York Regents Exams in Physical Setting Chemistry

CLEP Success Peterson's, 2007-05-04 Offers advice on improving scores on the CLEP and includes reviews for five CLEP subjects and five full length practice tests with explanatory answers

Master the CLEP Peterson's, 2012-04-30 Peterson's Master the CLEP is the essential prep guide for non traditional students who are looking to begin a college career and for current students who want to save on tuition costs by testing out of certain courses When you use this guide you will get a basic overview of the CLEP a look at the general examinations and the vast number of subject specific exams tips for registering and preparing for the CLEP and a guide to interpreting your test scores This is followed by extensive reviews of the five CLEP general examinations with pre and post tests and subject matter overviews that will help you pinpoint your strengths and weaknesses The Appendix will go into detail about a number of CLEP subject exams

Journal of Chemical Education, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers

The ETS Test Collection Catalog Educational Testing Service. Test Collection, 1993 The major source of information on the availability of standardized tests Wilson Library Bulletin Covers commercially available standardized tests and hard to locate research instruments

Optimizing STEM Education With Advanced ICTs and Simulations Levin, Ilya, Tsybulsky, Dina, 2017-06-05 The role of technology in educational settings has become increasingly prominent in recent years When utilized effectively these tools provide a higher quality of learning for students Optimizing STEM Education With Advanced ICTs and Simulations is an innovative reference source for the latest scholarly research on the integration of digital tools for enhanced STEM based learning environments

Highlighting a range of pivotal topics such as mobile games virtual labs and participatory simulations this publication is ideally designed for educators professionals academics and students seeking material on emerging educational technologies

Tests Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services

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Three Cognitive Skills in Chemistry and Their Application to Stoichiometry Ardas Ozsogomonyan, 1977 *A Summary of Science Education, 1988* Leopold E. Klopfer, 1990-06 **Nigerian Journal of Research and Production**, 2010-11

Golden State Examination, 1992 **Media Review**, 1986 *Cracking the AP Chemistry, 2002-2003 Edition* Paul Foglino, Princeton Review (Firm), 2002-01-29 The Princeton Review realizes that acing the AP Chemistry exam is very different from getting straight As in school The Princeton Review doesn t try to teach students everything there is to know about chemistry only the techniques they ll need to score higher on the exam There s a big difference In Cracking the AP Chemistry TPR will teach test takers how to think like the test makers and Score higher by reviewing key chemistry concepts Earn more points by becoming familiar with the format of the test Safeguard against traps that can lower scores Perfect skills with review questions in each chapter This book includes 2 full length simulated AP Chemistry exams All of The Princeton Review practice test questions are like the ones test takers will see on the actual exam and every solution is fully explained

Chemistry 2000-2001 Paul Foglino, 2000-01-25 The Book That Gets You Results Score higher by studying the most frequently tested material including atomic structure chemical equations kinetics thermodynamics and more We Know the AP Chemistry Exam The experts at The Princeton Review study the AP Chemistry exam and other standardized tests each

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