

# Chemistry: 2<sup>nd</sup> Semester Exam

Select the correct oxidation numbers for the following elements.

- Oxygen in  $O_2$ 
  - 2
  - +2
  - 1
  - +1
  - 0
- Sodium in  $NaCl$ 
  - 2
  - +2
  - 1
  - +1
  - 0
- Phosphorus in  $H_3PO_4$ 
  - 5
  - 3
  - 0
  - +3
  - +5
- H in  $H_2O$ 
  - 2
  - +2
  - 1
  - +1
  - 0
- O in  $H_2O$ 
  - 2
  - +2
  - 1
  - +1
  - 0
- In which substance does bromine have an oxidation number of +1?
  - $Br_2$
  - $HBr$
  - $HBrO$
  - $HBrO_2$
- Which species is the oxidizing agent in the following reaction?  $Cl_2(aq) + 2I^-(aq) \rightarrow I_2(aq) + 2Cl^-(aq)$ 
  - $Cl_2$
  - $2I^-$
  - $I_2$
  - $2Cl^-$
- Consider the following:  $2MnO_4^- + 5H_2SO_3 \rightarrow 2Mn^{2+} + 3H_2O + 5SO_4^{2-} + 4H^+$   
Which of the following elements does not experience a change in oxidation number?
  - S in  $H_2SO_3$
  - O in  $MnO_4^-$
  - Mn in  $MnO_4^-$
- In which of the following unbalanced equations does chromium undergo oxidation?
  - $Cr^{3+} \rightarrow Cr$
  - $Cr^{2+} \rightarrow Cr^{3+}$
  - $Cr^{3+} \rightarrow Cr_2O_7^{2-}$
  - $CrO_4^{2-} \rightarrow Cr_2O_7^{2-}$
- In a redox reaction, the species which loses electrons is \_\_\_\_\_.
  - reduced
  - oxidized
  - neutralized
- Consider the following:  $Cl_2 \rightarrow ClO_2^-$   
The balanced half reaction is:
  - $Cl_2 + 3H_2O \rightarrow 2ClO_2^- + 6H^+ + 5e^-$
  - $Cl_2 + 3H_2O \rightarrow 2ClO_2^- + 6H^+ + 4e^-$
  - $Cl_2 + 6H_2O + 2e^- \rightarrow 2ClO_2^- + 6H_2$
  - $Cl_2 + 6H_2O \rightarrow 2ClO_2^- + 12H^+ + 10e^-$

## Label the Acid, Base, Conjugate Acid, and Conjugate Base



- $HC_2H_3O_2$  serves as the
  - Acid
  - Base
  - Conjugate Acid
  - Conjugate Base
- $H_2O$  serves as the
  - Acid
  - Base
  - Conjugate Acid
  - Conjugate Base
- $H_3O^+$  serves as the
  - Acid
  - Base
  - Conjugate Acid
  - Conjugate Base
- $C_2H_3O_2^-$  serves as the
  - Acid
  - Base
  - Conjugate Acid
  - Conjugate Base

# 2nd Semester Exam Chemistry Form

**William D. Hedges**



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