

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide ion concentrations, molar ratios, and pH values, and vice versa, for weak acid solutions (at 25°C).
3. Calculate the pH of weak (or strong) acids.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydroxide ion | 5. molar ratio | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|--|
| 1. $\text{pH} = -\log_{10} [\text{H}^+]$ (Eq. 14.1) | 2. $\text{pOH} = -\log_{10} [\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3) | 4. $\text{pH} = \text{pK}_a + \log_{10} \frac{[\text{A}^-]}{[\text{HA}]}$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number of significant figures in the answers to the questions below. Round answers to the correct number of significant figures.

1. What products are **1** formed when hydrochloric acid (HCl) ionizes in water? The concentration of these acids produced in 0.100 M HCl solution is **0.100** M.
2. Hydrochloric acid ionizes from **1** to **0.100** M. How many **0.100** moles of H^+ ions are produced in 1.00 L of 0.100 M HCl solution? **0.100** moles.
3. How many **0.100** moles of H^+ ions are produced in 1.00 L of 0.100 M HCl solution? **0.100** moles.
4. How many **0.100** moles of H^+ ions are produced in 1.00 L of 0.100 M HCl solution? **0.100** moles.
5. How many **0.100** moles of H^+ ions are produced in 1.00 L of 0.100 M HCl solution? **0.100** moles.

6. **0.100** M
7. **0.100** M
8. **0.100** M
9. **0.100** M
10. **0.100** M
11. **0.100** M
12. **0.100** M
13. **0.100** M
14. **0.100** M
15. **0.100** M
16. **0.100** M

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