



# 70 S Show Episode Guide

**S Marginson**



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web write down the equilibrium equation you are investigating using the information provided by the introduction within the virtual lab activity equilibrium reaction  $CO_2 + H_2O \rightleftharpoons H_2CO_3$  select each control test tube to compare the concentrations of

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web equilibrium moles hscn  $2.00 \times 10^{-5}$  mol  $3.00 \times 10^{-6}$  mol  $1.70 \times 10^{-5}$  mol hscn knowing the number of moles of  $Fe^{3+}$  and hscn present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species mol  $Fe^{3+}$   $3.170 \times 10^{-5}$  mol  $Fe^{3+}$   $3.850 \times 10^{-4}$  m

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