

Page 4	Mark Scheme	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2013	9702	22

7	(a)	(i)	$W = 206$ and $X = 82$ $Y = 4$ and $Z = 2$	A1 A1	[2]
		(ii)	mass-energy is conserved mass on rhs is less because energy is released	B1 B1	[2]
	(b)	not affected by external conditions/factors/environment or two examples temperature and pressure	B1	[1]	

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less

exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with hcl as  $57.9 \text{ kJ mol}^{-1}$

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example  $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(aq)}$  you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts