



1997 Ford E 350 73 Repair Manual

M Woodhall



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 $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
 for every two molecules of ammonia made four molecules of reactant are used up this equilibrium shifts to the right with an increase in pressure

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 $\text{CO}(\text{g}) + 2\text{H}_2(\text{g}) \rightleftharpoons \text{CH}_3\text{OH}(\text{g})$

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