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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$   
oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$  reduction  $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni}^{3+} \rightarrow \text{Ni} + 3\text{Ag}$  oxidation reduction 25  $\text{Cu} + \text{Pb}^{2+} \rightarrow \text{Pb} + \text{Cu}$  oxidation reduction 26  $\text{O}_2 + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O}$

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 $5\text{CO} + \text{I}_2 \rightarrow 5\text{CO}_2 + \text{I}_2$   $\text{O}_5 + 5\text{C} \rightarrow \text{I}_2 + 5\text{CO}_2$

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state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr}^0$  to  $\text{Sr}^2$  oxidized reducing agent  $\text{O}^0$  to  $\text{O}^2$  reduced ox ag  $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li}^0$  to  $\text{Li}^1$  oxidized red ag  $\text{S}^0$  to  $\text{S}^2$  reduced ox ag  $3\text{C} + 2\text{Br}_2 \rightarrow \text{C}_2\text{Br}_4$