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Lectures, Speeches, and
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ROBERT GREEN INGERSOLL

A
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BOOK

Ingersollia Gems Of Thought From The Lec

WJ Hussar



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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2 \text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$ H_2O $2 \text{Cr}^{3+} \text{aq}$ $2 \text{OH}^- \text{aq}$

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web jan 29 2020 the ionic form of this reaction is $2 \text{Na}^+ \text{aq}$ $2 \text{Cl}^- \text{aq}$ $\text{Cu}^{2+} \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $2 \text{Na}^+ \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $\text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3 \text{Ba}^{2+} \text{aq}$ $2 \text{PO}_4^{3-} \text{aq}$ $\text{Ba}_3\text{PO}_4\text{2 s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3 \text{aq} + \text{KCl aq} \rightarrow \text{AgCl s} + \text{KNO}_3 \text{aq}$ total ionic $\text{Ag}^+ \text{aq} + \text{NO}_3^- \text{aq} + \text{K}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl s} + \text{K}^+ \text{aq} + \text{NO}_3^- \text{aq}$ net ionic $\text{Ag}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl s}$

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oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O l}$ the spectator ions in $\text{KOH aq} + \text{HNO}_3 \text{aq} \rightarrow \text{K}^+ \text{aq} + \text{NO}_3^- \text{aq}$ how to identify and list the

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web the Ag^+ and Cl^- ions become AgCl s but the Na^+ ions and the NO_3^- ions stay as Na^+ ions and NO_3^- ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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