

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous acids/bases.
2. Calculate hydroxide and hydronium ion concentrations and pH/pOH and the associated logarithmic functions.
3. Calculate the pH/pOH of an aqueous solution.

## Vocabulary

- |                   |                                |                 |
|-------------------|--------------------------------|-----------------|
| 1. autoionization | 2. conjugate acid-base pair    | 3. pH           |
| 4. hydronium ion  | 5. hydronium ion concentration | 6. pOH          |
| 7. hydroxide ion  | 8. hydroxide ion concentration | 9. $\text{pOH}$ |

## Key Equations

- |   |   |
|---|---|
| 1. $\text{pH} = -\log_{10} [\text{H}_3\text{O}^+]$ (Eq. 14.1) | 2. $\text{pOH} = -\log_{10} [\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3)          | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.4)        |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair associated with the species and indicate which member of the pair is the stronger acid.

1.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
2.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
3.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
4.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
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8.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
9.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
10.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
11.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
12.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
13.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
14.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
15.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
16.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
17.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
18.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
19.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
20.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )

## Part B: Multiple Choice

1.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
2.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
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10.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
11.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
12.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
13.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
14.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
15.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
16.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
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18.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )
19.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{H}_2\text{SO}_4$ )
20.  $\text{H}_2\text{SO}_4$  and  $\text{HSO}_4^-$  (stronger acid is  $\text{HSO}_4^-$ )

# 19 2 Hydrogen Ions And Acidity Section Review

**ML Morrison**



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