



# Aidan Chambers Of Cops Robbers

**Clemens Wendtner**



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web answer a 1 explanation the entire reaction was multiplied by 2 therefore  $\Delta H$  must also be halved 3 answer c explanation option c is incorrect because 566 kJ mol is the  $\Delta H$  per mole of o<sub>2</sub> and not for 2 moles even though options b and d are expressed in grams they are still correct because 56

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web  $q = 140 \text{ g} \times 4.184 \text{ J} / \text{g} \cdot \text{C} \times (16.0 \text{ C} - 93.72 \text{ C}) = 5453 \text{ J}$  1 mole  $\text{KClO}_3$  5453 moles  $\text{H}_2\text{O}$  5453 mole moles 5 when a 15 g sample of  $\text{SrCl}_2$  dissolves in 80 g of water in a calorimeter the temperature drops from 42 C to 15 C calculate the  $\Delta H$  for the process

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