

## 14.5

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                             |                   |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid      |
| 4. autoionization | 5. hydronium ion            | 6. autoionization |
|                   | 7. weak base                | 8. pH             |

## Key Equations

- |  |  |
|--|--|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (14.1) | 2. $\text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) |
| 3. $\text{pH} = -\log [\text{H}^+]$ (14.3)                                 | 4. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ (14.4)                                 |

## Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate species.

1. What products are \_\_\_\_\_ between hydroxide ion and hydronium ion? \_\_\_\_\_ the conjugate base of the species \_\_\_\_\_ produced in this reaction.
2. Hydroxide ion and hydronium ion \_\_\_\_\_ in water.
3. \_\_\_\_\_ is the conjugate base of the \_\_\_\_\_ ion.
4. \_\_\_\_\_ is the conjugate \_\_\_\_\_ of \_\_\_\_\_.
5. \_\_\_\_\_ is the conjugate base of \_\_\_\_\_.
6. \_\_\_\_\_ is the conjugate \_\_\_\_\_ of \_\_\_\_\_.
7. \_\_\_\_\_ is the conjugate base of \_\_\_\_\_.
8. \_\_\_\_\_ is the conjugate \_\_\_\_\_ of \_\_\_\_\_.

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# 19 2 Hydrogen Ions And Acidity Answer Key

**SJ Ball**



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