

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl **hydrochloric acid**
- H₂SO₄ **sulfuric acid**
- HNO₃ **nitric acid**
- HClO₄ **perchloric acid**
- Strontium hydroxide **LiOH**
- potassium hydroxide **KOH**
- calcium hydroxide **CaOH**
- calcium hydroxide **Ca(OH)₂**
- ammonium hydroxide **Li(OH)₄**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.
HClO₄ is STRONG with a lower cell completely dissociate in water

HCl **H₂** **H₂O** **Mg(OH)₂** **HNO₃** **Ca(OH)₂** **HClO₄** **H₂SO₄** **H₂CO₃**

3. What is the formula for hydrosulfuric acid? **H₂S**

4. MgCO_3 will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction. **$\text{MgCO}_3 + \text{H}_2\text{O} \rightleftharpoons \text{Mg}^{2+} + \text{HCO}_3^-$**

b. Would you expect a large or small equilibrium constant for this reaction? **small, because MgCO_3 is a weak acid.**

c. Would you define MgCO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**

d. In your written equation, is HCO_3^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**

e. Identify the conjugate base of MgCO_3 . **HCO_3^-**

f. Identify MgCO_3 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's.**

5. F^- will react with water and gain a H^+ ion in the process.

a. Write a balanced equation for this reaction. **$\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$**

b. Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**

c. Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**

d. In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**

e. Identify the conjugate acid of F^- . **HF**

f. Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HCO_3^- will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction. **$\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{CO}_3^{2-}$**

b. Would you expect a large or small equilibrium constant for this reaction? **small, because HCO_3^- is a weak acid.**

c. Would you define HCO_3^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**

d. In your written equation, is HCO_3^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**

e. Identify the conjugate base of HCO_3^- . **CO_3^{2-}**

f. Identify HCO_3^- as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only one H's) and H's (can't lose all H's in one go)**

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Simeon Shaw



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