

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. autoionization |
| | 7. weak acid | 8. pH |

Key Equations

- | | |
|--|------------------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (Eq. 17.1) | 2. $pH = -\log[H_3O^+]$ (Eq. 17.2) |
| 3. $pH = -\log[H_3O^+]$ (Eq. 17.3) | 4. $pH = -\log[H_3O^+]$ (Eq. 17.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Write the number in the space provided.

1. H_2O and H_3O^+ are _____, whereas H_2O and OH^- are _____.
2. H_2O and OH^- are _____, whereas H_2O and H_3O^+ are _____.
3. H_2O and H_3O^+ are _____, whereas H_2O and OH^- are _____.
4. H_2O and OH^- are _____, whereas H_2O and H_3O^+ are _____.
5. H_2O and H_3O^+ are _____, whereas H_2O and OH^- are _____.
6. H_2O and OH^- are _____, whereas H_2O and H_3O^+ are _____.
7. H_2O and H_3O^+ are _____, whereas H_2O and OH^- are _____.
8. H_2O and OH^- are _____, whereas H_2O and H_3O^+ are _____.

9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. _____

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