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web we then changed the volume between 5ml and 20ml and recorded the pressure at each this part of the lab showed that volume and pressure are inversely proportionate as the pressure would decrease if the volume was increased part b of the experiment looked at the relationship between temperature and pressure

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web procedure 1 the relationship between pressure and volume set the plunger of the syringe to 1 10 the capacity of the syringe the 2 0 ml mark if using a 20 ml syringe the volume is read from the edge of the bottom rib of the plunger attach the syringe to the connector at the end of sensor

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1/267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1/273 and concluded that the pressure would approach zero at 1/273 degrees c

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web the volume of 1.00 mol of any gas at stp standard temperature 273.15 k and pressure 1 atm is measured to be 22.414 l we can substitute 101.325 kpa for pressure 22.414 l for volume and 273.15 k for temperature into the ideal gas equation and solve for r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of gases 273.15 k and 1 atm 101.325 kpa 1 at stp one mole of an ideal gas has a volume of about 22.4 l this is referred to as the standard molar volume figure 9.18

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web this lab explores the relationships between pressure volume and temperature in an ideal gas the ideal gas law  $pV = nRT$  is used and with it we can explore it by seeing how the pressure and volumes changed in the tubes each time that we

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web pressure increases even though there were errors in the experiment the goal of the experiment was met the goal of declaring what type of relationship it is was clear using the data pressure kpa temperature c temperature k constant k p t 95.82 kpa 16.0 c 289 k 0.33 97.90 kpa 19.0 c 292 k 0.34 98.97 kpa 99.22 kpa 101.4 kpa 19.1 c

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