

Chapter 9: Chemical Quantities

1. Solve stoichiometry problems using grams \rightarrow moles \rightarrow moles \rightarrow grams calculations. Examples below:

a. Calculate moles/grams of $(\text{NH}_4)_2\text{SO}_4$ produced when 30.0 moles of NH_3 react with H_2SO_4 .



30.0 mol x moles

$$30.0 \text{ moles} \times 1 \text{ mol}/2 \text{ mol} = 15.0 \text{ moles}$$

$$15.0 \text{ moles } (\text{NH}_4)_2\text{SO}_4 \times 132.2 \text{ grams/mole} = 1983 = 1980 \text{ grams (to 3 s.f.)}$$

b. $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$. How many moles/grams of sodium nitrate are produced by the complete reaction of 100. grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? *Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.*

Grams \rightarrow Moles \rightarrow Moles \rightarrow Grams

$$100. \text{ g Na}_2\text{CO}_3 \times 1 \text{ mol}/106.0 \text{ g} \times 2 \text{ mol NaNO}_3 / 1 \text{ mol Na}_2\text{CO}_3 \times 85.00 \text{ g/mol} = 160.4 \text{ g}$$

$$500.0 \text{ g NaNO}_3 \times 1 \text{ mol}/85.00 \text{ g} \times 2 \text{ mol} / 2 \text{ mol} \times 63.02 \text{ g/mol} = 370.7 \text{ grams}$$

2. Calculate percent yield of products. Example below using (b) above.

What is the theoretical yield of NaNO_3 if 100.0 grams of Na_2CO_3 are used in the reaction? What is the percent yield of NaNO_3 if 130.5 grams of NaNO_3 are actually collected?

Theoretical yield = 160.4 grams (answer from above)

$$\text{Percent Yield} = (\text{Actual Yield}/\text{Theoretical Yield}) \times 100 = (130.5 \text{ g}/160.4 \text{ g}) \times 100 = 81.36\%$$

3. What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

Chapter 10: Modern Atomic Theory

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

5. Define electromagnetic radiation, wavelength, and frequency.

6. How do atoms gain energy to become excited? *Absorb a photon of energy.*

How do they lose energy to return to their ground state? *Release a photon of same energy as that absorbed.*

7. What is the evidence that energy absorbed by the hydrogen atom is quantized? *Specific wavelengths of light seen in emission spectra for atoms (page 285-86)*

8. Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, probs 54, 96

2nd Semester Exam Chemistry Form

**Jay Wackerly, Sarah Zingales, Michael
Wentzel, Gautam Bhattacharyya, Brett
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University of Iowa Studies in Education, 1930 **University of Iowa Studies in Education** George Dinsmore Stoddard, University of Iowa, 1925 **Official Gazette** Philippines, 2004 **Organic Chemistry Education Research into Practice** Jay Wackerly, Sarah Zingales, Michael Wentzel, Gautam Bhattacharyya, Brett McCollum, 2025-03-25 This Research Topic has three main goals 1 provide a platform for instructors of organic chemistry to showcase evidence based methods and educational theories they have utilized in their classrooms 2 build new and strengthen existing connections between educational researchers and practitioners and 3 highlight how people have used chemical education based research in their teaching practice There are places in the literature dedicated for chemical education research CER however there is not a clear avenue for those that have changed their teaching methods based on published CER and report their experiences Creating this article collection will foster collaboration between chemical education researchers and teachers of organic chemistry This opportunity allows these instructors to share evidence based practices experiences challenges and innovative approaches from CER literature and beyond This Research Topic bridges discipline based education research and the scholarship of teaching and learning which will help advance organic chemistry education and improve student outcomes

An Experimental Study of the Iowa Placement Examinations Lawrence Williams Miller, University of Iowa, 1930
Circular of Information USA. Bureau of Education, 1886 **Oxford University Gazette** University of Oxford, 1916
Program Bryn Mawr College, 1905 British Medical Journal, 1913 **Contagious Diseases of domesticated Animals** D. E. Salmon, 1884 **Technical Education, an Economic Necessity** Victor Clifton Alderson, 1885 **The Universal tutor** London gen. corresp. coll, 1894 **The Lancet**, 1888 **The Prospective Teacher's Grasp of High School Content in His Probable Fields of Instruction** Russell E. Jonas, University of Iowa, 1928 Testing and Evaluation for the Sciences in the Secondary School William D. Hedges, 1966 *Bryn Mawr College Calendar* Bryn Mawr College, 1924 *A Project in Student Personnel Service Designed to Facilitate Each Student's Achievement at the Level of His Ability* Lonzo Jones, 1930 Calendar Bryn Mawr College, 1885 *School Science and Mathematics*, 1921 **Science Tests and Reviews** Buros Center, 1975 Science Tests and Reviews consisting of science sections of the first seven MMYs and Tests in Print II includes 217 original test reviews written by 81 specialists 18 excerpted test reviews 270 references on the construction use and validity of specific tests a bibliography on in print science tests references for specific tests cumulative name indexes for specific tests with references a publishers directory title index name index and a scanning index The 97 tests covered fall into the following categories 23 general 14 biology 35 chemistry 3 geology 6 miscellaneous and 16 physics

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