

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. autoionization | 5. hydronium ion | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|--|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (17.1) | 2. $\text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (17.2) |
| 3. $\text{pH} = -\log[\text{H}^+]$ (17.3) | 4. $\text{pOH} = -\log[\text{OH}^-]$ (17.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the _____ of water's autoionization?
2. _____ is the conjugate base of the autoionization of water which produces an acidic solution.
3. _____ is the conjugate acid of the autoionization of water which produces a basic solution.
4. _____ is the conjugate acid of the autoionization of water which produces a neutral solution.
5. _____ is the conjugate base of the autoionization of water which produces a neutral solution.
6. _____ is the conjugate acid of the autoionization of water which produces an acidic solution.
7. _____ is the conjugate base of the autoionization of water which produces an acidic solution.

8. _____
9. _____
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15. _____
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17. _____
18. _____

19 2 Hydrogen Ions And Acidity Answers

Wisconsin Dairy and Food Commission



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