

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous acids/bases.
2. Calculate K_a and K_b concentrations of acids/bases and K_a/K_b and K_b/K_a of conjugate acid-base pairs.
3. Calculate pH and pOH of solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|---------------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. K_a and K_b | 6. pH and pOH |

Key Equations

- | | |
|---|---|
| 1. $\text{pH} = -\log[\text{H}^+]$ (Eq. 14.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ | 4. $\text{pH} + \text{pOH} = 14.00$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the _____ of a weak hydroxide ion and a hydrochloric acid? The conjugate base of this acid is produced in this reaction. _____
2. Hydrochloric acid ionizes from _____ to form _____.
3. _____ ionizes to form _____ and _____.
4. _____ ionizes to form _____ and _____.
5. _____ ionizes to form _____ and _____.
6. _____ ionizes to form _____ and _____.
7. _____ ionizes to form _____ and _____.
8. _____ ionizes to form _____ and _____.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Jiyuan Zhang



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